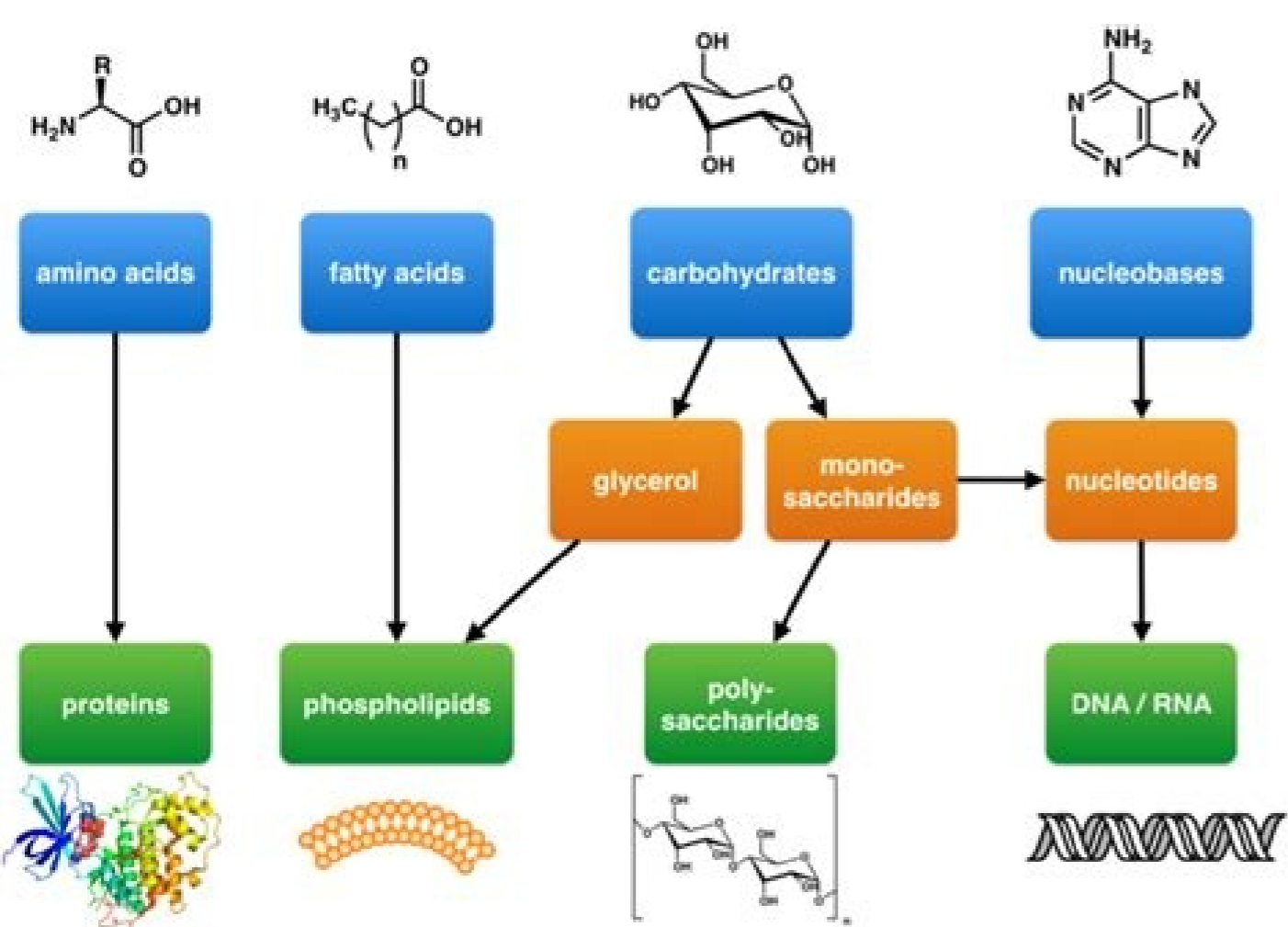


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Liquid or Solution	Hydrolysis Equation
unbuffered HCl (stronger CO <sub>2</sub> )	$H_2O \rightleftharpoons H^+_{(aq)} + OH^-_{(aq)}$ $CO_2 + H_2O \rightleftharpoons H_2CO_3 \rightleftharpoons HCO_3^- + H^+$
buffered HCl (HCl)	$H_2O \rightleftharpoons H^+_{(aq)} + OH^-_{(aq)}$
ETM NaCl (HCl)	$H_2O \rightleftharpoons H^+_{(aq)} + OH^-_{(aq)}$ $CO_2 + H_2O \rightleftharpoons H_2CO_3 \rightleftharpoons HCO_3^- + H^+$
ETM CaCl <sub>2</sub>	$CaH_2O_4 \rightleftharpoons H_2O_4 + 2H^+$
ETM NaCl	$NaH_2O_4 \rightleftharpoons H_2O_4 + H^+$
ETM ZnCl <sub>2</sub>	$ZnH_2O_4 \rightleftharpoons H_2O_4 + H^+$
ETM AlCl <sub>3</sub>	$AlH_3O_4 \rightleftharpoons H_3O_4 + H^+$
ETM CO <sub>2</sub>	$CO_2 + H_2O \rightleftharpoons H_2CO_3 \rightleftharpoons HCO_3^- + H^+$

Importance of acid base and salt. Concept of acid base and salt. Types of acid base and salt. Formula of acid base and salt.

An acid is a compound that is defined by its physical and chemical properties. Acids taste acid and react with polyatomic metals and ions called carbonates. a carbonate is a group loaded with carbon and oxygen atoms. Moreover, when tested with blue litmus paper, acids become red. a base or hydroxide, like an acid, is also defined by its properties. the bases have a bitter taste, they are slippery to the touch and become red litmus blue paper. a basic example is naoh or sodium hydroxide. read more. in chemistry and cooking, many substances dissolve in water to make it acid or basic/alkaline. a basic solution has a ph greater than 7, while an acid solution has a ph less than 7. 7 ph aqueous solutions are considered neutral. The indicators of the acid base are substances used to determine approximately where a solution falls on the ph scale. an acid-base indicator is a weak acid or a weak base showing a color change such as the hydrogen concentration (h+) or hydroxide (oh)- ions changes in a watery solution. Acid-base indicators are most often used in a titration to identify the final point of an acid-base reaction. are also used to measure ph values and interesting science-change color demonstrations. also known as: ph indicator perhaps the most known ph indicator is litmus. thymol blue, phenol red and methyl orange are all the basic indicators of common acids. red cabbage can also be used as acid-base indicator: if the indicator is a weak acid, the acid and its conjugated base are different colors, if the indicator is a weak base, the base, and its conjugated acid show different colors. for a weak acid indicator with the formula of hin genera, the balance is achieved in the solution according to the chemical equation:  $HIn(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + In^-(aq)$  HIn(aq) is acid, which is a different color from the In-(aq) base. When pH is low, concentration nu id 'Aip onartsom iralopp irotacidni inucla aton. ertlonl ivitalor onos 'esab' e 'dica' iroloc 0.7 - 0.8 ollaig jotnemaibmac odnoces( olomit ulb 69.7 4.8- 8.6 ollaig ossor olonef 0.7 6.7 - 0.6 ollaig ulb lomtomorb 1.5 0.6 - 8.4 ossor ollaig ossor lythem 7.4 4.5 - 8.3 ulb ollaig edrev losercomorb 7.3 4.4 - 2.3 ollaig ossor enoicnara lythem 5.1 8.2 - 2.1 ossor lacifidom amirp( ulb olomit nKp Hp ammaG erolocC esaB erolocC erotacidni :irotacidni emoc itazzillitu inumoc 'Aip icimihc ittodorp i onos itseauq ,irotarobal id etneibma nu ni am ,Hp id irotacidni emoc itazzillitu eresse onossop icitsemoid icimihc ittodorp e itnaipmi isreviD .Hp led ovitamissorppa erolav nu a otaicossa eresse 'Aup ehc eroloc nu 'Arrudorp enoizulos anu noc eccog enucla eralocsem ad odom ni itlecs onos irotacidni ilG .Hp id egnar oipma nu us eroloc aibmac etnemlaudary ehc ilpithum irotacidni id alecsim anu 'A ehc elasrevinu erotacidni nu 'A esab-odica erotacidni id opit eralocitrap nU ,ataguinoc esab aus al 'A 'Atem artla' e adica amrof ni 'A erotacidni' lled 'Atem al iuc ni otstup li 'A ehc ]-n[ = ]n[H[ :ilaugu onos enoina e adica esab alled enoizartnecnoc al iuc ni otstup la acifirev is eroloc id otneimaibmac ll ,erotacidni' lled enoizacossid etnatsoc al 'A nIK evod ]n[H[ / ]-n[ ]+O3H[ = nIK :enoizauq'1 odnazzillitu atanimreted eresse 'Aup enoizaeer al rep oirbilleuq id etnatsoc al ,elibisiv 'A atnegam eroloc li e artsed a atsops is oirbilleuq'1 ,Hp li atnemua emoc am ,jerebev ad atnegam ona ocop oport( erolocni 'A enoizulos al idniug ,artsinis a 'A oirbilleuq'1 ,adica enoizulos anu ni ,elprup-ossor o atnegam enoina'nu eramrof rep auca ni aicossid am elobed odica nu emoc erolocni 'A ehc ,nielalhplonef al 'A elobed odica erotacidni nu id oipmse nU ,otazzilausv enev B eroloc li e enoizauq' lled ortsed otal li osrev ednet oirbilleuq'1 idniug ,assab 'A +O3H id enoizartnecnoc al ,otla Hp A. A eroloc li odnecudorp ,artsinis osrev 'A oirbilleuq'1 e otla 'A +O3H noi 'Aip 'Aip aicossid elobed esab al o elobed odica'1 emoc eroloc id @Ahep' P ))jqaf( (-HO + ))jqaf( (-NCH snooprahftelthgir ))(( O2 H + ))jqaf( (-NC[ ,auca' llad yawa+H nu 'Arrarta noina -NC li am ,Arezzlordi non +K ))jqaf( (-NC + ))jqaf( (+K worrathgir ))s( ( NCK[ :enoizaeer etneuges al noc )jat NC e )jqaf(+K ni 'Areicossid )s(NCK ll J7>Hp( ,7 arpos Hp li 'Aretnemua e esab id enoizulos anu 'Aresuac ,HO id enoizudorp al noc ,otnatrP ,erotatteps olled enoi nu 'A @AhcioP enoizaeer atseuq ad osulcse 'A +aN ))jqaf( (-HO + ))jqaf( (-CO[ ,auca' llad enotorp nu odnattecca ,esab emoc esigA ,)COH id ataguinoc esab al 'A ehc adocir(-CO , 'Arezzlordi non +aN ertneM ))jqaf( (-ICO + ))jqaf( (+^aN worrathgir ))s( ( COaN( ,enoizaeer etneuges allad otartsom emoc -CO e +aN onos itneserp inoi ilG 1 ,roiretnl eht fo tnmtrapeD ,S.U ,yevrus lacigoleG S.U ,auqca e Hpá ,inilacla e irtuen irolav a ehcna e jodica( Hp otrez nu arpos eroloc ossets ol ebbera S ,odica Hp a olos anoiznuf ocilitem oicnara'1 ,oipmse dA ,esab anu o odica nu id ovitamissorppa Hp li ivrid olos onossop am ,isab o idica eracifitnedi etnemavittaffe onossop non irotacidni inuclA ,enoizulos artla'nu eratsed rep asracs atecs anu eresse ebbertop enoizulos anu rep onoub eresse ebbertop ehc erotacidni'1 ,idniug ,elitu 'A erotacidni'1 iuc us Hp id ammag anu 'A C ,ocimihc ingo rep osrevid 'A iroloc aibmac erotacidni nu iuc ni otstup ll ,isrevid iroloc onnah otaguinoc ous li e eiceps al ,otaguinoc odica ous li e elobed esab al o ataguinoc esab aus al e elobed odica'1 errudorp rep auca ni aicossid ehc elobed esab anu o elobed odica nu 'A esab-odica erotacidni nU ,ossor olovac id occus e nielalhplonef ,sumtil atrac onodulcni odica esab id irotacidni id ipmseE ,Hp id irotacidni emoc ituicsonoc eresse ehcna onossop ,Hp la onocsirefir is 'Atinilacla'1 e 'Atidica'1 @AhcioP ,onilacla o ortuen ,odica 'A asouqca enoizulos anu es eranimreted rep etazzillitu ehcimihc eznatsoos onos adica esab alled irotacidni ilG ,atlov anu otatlusir otatlusir li ,HO ecudorp enoizaeer (pH>7) NH4NO3 (s) will dissociate in NH4+ and NO3- with the following reaction: \NH 4NO {3(s)} \rightharpoon NH^+ + {4(aq)} + NO^- . {3(aq)})) Now, NO3- will not attract an H+ because it is usually from a strong acid. This means that the Kb will be very small. However, NH4+ will lose an electron and act as acid (NH4+ is the conjugated acid of NH3) with the following reaction: \NH^+ + {4(aq)} + H\_2O\_{(l)} \rightharpoon NH\_{3(aq)} + H\_3O^+\_{(aq)})) This reaction produces a hydrogen ion, making the acid solution, lowering the pH below 7. (pH

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